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IN RE APPLICATION OF:

Sebastian C. Reyes, et al.

SERIAL NUMBER: 10/699,258

FILED: October 31, 2003

**FOR: LIGHT HYDROCARBON
SEPARATION USING
8-MEMBER RING ZEOLITES**

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EXAMINER: Tam. M. Nguyen

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**Commissioner for Patents
P.O. Box 1450
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DECLARATION UNDER 37 CFR §1.132

1. I, Sebastian C. Reyes, do declare and say as follows:
2. I am a co-inventor of the invention set forth in the above-noted patent application and I am making this declaration in support of the patent application;
3. I am presently employed with ExxonMobil Research and Engineering Company;
4. I am a Senior Research Associate and have been working in the field of petroleum fuels processes for 21 years, providing research guidance and engineering;
5. My academic training is set forth as follows:
 - University of Minnesota, Ph.D. Chemical Engineering, 1985
 - University of Concepción, Chile, M.S. Chemical Engineering, 1981
 - University of Concepción, Chile, B. S. Chemical Engineering, 1979

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Reply to Office Action of August 22, 2006

6. I am familiar with the disclosure of the subject matter in the subject patent application;

7. I have read the Examiner's Office Action mailed August 22, 2006 concerning the subject patent application and claim rejections made therein;

8. In my opinion, to one of ordinary skill in the art, the use of the term " $\text{SiO}_2/\text{Al}_2\text{O}_3$ ratio" as used in the present application is equivalent to " $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio". In either industry or academy, in the context of defining a zeolite framework composition, the two terms are used interchangeably. When using the term " $\text{SiO}_2/\text{Al}_2\text{O}_3$ ratio", there is never a confusion that such a ratio would refer to a mass basis. Only a molar basis has a proper physical meaning attached to it.

Perhaps it is instructive to point out that the instances in which either of the terms " $\text{SiO}_2/\text{Al}_2\text{O}_3$ mass ratio" or "mass $\text{SiO}_2/\text{Al}_2\text{O}_3$ ratio" appear in writing in either the patent or journal literature are practically zero relative to the term " $\text{SiO}_2/\text{Al}_2\text{O}_3$ ratio". A quick web search returns zero hits for the former terms but abundant hits for the latter.

A clear illustration that the term " $\text{SiO}_2/\text{Al}_2\text{O}_3$ ratio" is commonly used without ambiguity, comes from a classical and respected book in the area of adsorptive separations. The book is called "Principles of Adsorption and Separation Processes" (D. M. Ruthven, John Wiley and Sons, New York, 1984), whose contents are very pertinent to the subject matter of the present application. On page 10 of this reference (copy of page attached), in the very first instance in the book in which the author refers to framework composition in zeolites states: "The Si/Al ratio in the zeolite...". Such wording is ubiquitous throughout the book but there no confusion that the implied meaning is a molar ratio.

I further declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed true; and that these statements were made with the knowledge that willful false statements and the like so

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made are punishable by fine or imprisonment, or both, under Section 1001 of Title 18 of the United States Code, and that such willful false statements may jeopardize the validity of the application or any patents issuing thereon.

Signed: Sebastian C. Reyes
Sebastian C. Reyes

Date: Dec/6/2006

PRINCIPLES OF ADSORPTION AND ADSORPTION PROCESSES

DOUGLAS M. RUTHVEN

University of New Brunswick, Fredericton

A Wiley-Interscience Publication

JOHN WILEY & SONS

New York Chichester Brisbane Toronto Singapore

10 Microporous Adsorbents

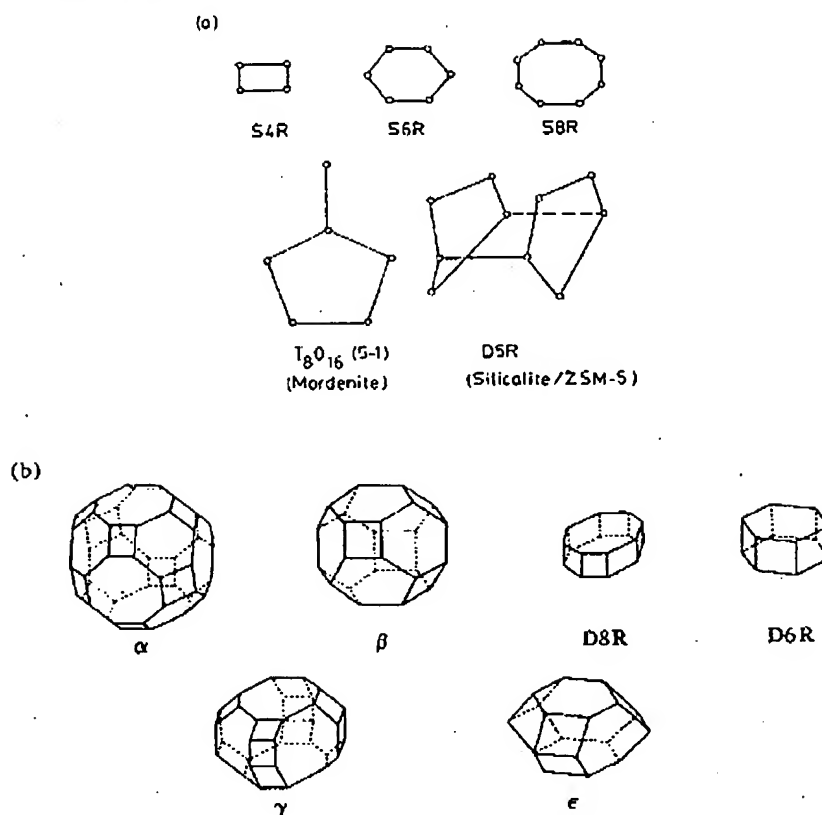


FIGURE 1.4. (a) Secondary building units and (b) commonly occurring polyhedral units in zeolite framework structures. (From ref. 19, copyright John Wiley & Sons, Inc., 1974; reprinted with permission.)

locations has been recently summarized by Mortier.⁽²⁵⁾ Changing the exchangeable cation by ion exchange provides a useful and widely exploited means of modifying the adsorptive properties.

The Si/Al ratio in a zeolite is never less than 1.0 but there is no upper limit and pure silica analogs of some of the zeolite structures have been prepared. The adsorptive properties show a systematic transition from the aluminum-rich sieves, which have very high affinities for water and other polar molecules, to the microporous silicas such as silicalite which are essentially hydrophobic and adsorb *n*-paraffins in preference to water. The transition from hydrophilic to hydrophobic normally occurs at a Si/Al ratio of between 8 and 10. By appropriate choice of framework structure, Si/Al ratio and cationic

NaX

CaX

FIGURE 1
Lennard-Je
occurs betw
Sons, Inc.,